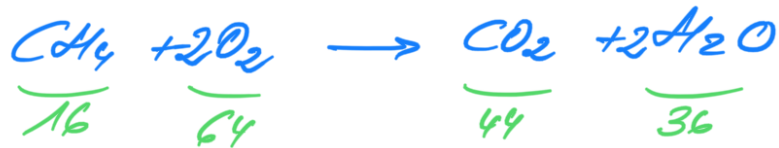


EJERCICIO FQ4EE2501:

COMBUSTIÓN DEL METANO (CH₄)

¿g H₂O? con 5 litros CH₄ a 20°C y 600 mmHg



$$V = 5 \text{ litros de CH}_4 \quad P = \frac{600}{760} \text{ atm}; \quad T = 20 + 273$$

$$P \cdot V = n \cdot R \cdot T$$

$$\frac{600}{760} \cdot 5 = \frac{g_{\text{CH}_4}}{16} \cdot 0,082 \cdot (20 + 273)$$

$$g_{\text{CH}_4} = 2,63 \text{ g}$$

$$2,63 \text{ g CH}_4 \cdot \frac{36 \text{ g H}_2\text{O}}{16 \text{ g CH}_4} = 5,92 \text{ g H}_2\text{O}$$